

# Section Review

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## Electrons and Chemical Bonding

### USING KEY TERMS

1. Use the following terms in the same sentence: *chemical bond* and *valence electron*.

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### UNDERSTANDING KEY IDEAS

- \_\_\_\_\_ 2. Which of the following atoms do not usually form bonds?  
a. calcium  
b. neon  
c. hydrogen  
d. oxygen

3. Describe chemical bonding.

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4. Explain how to use the valence electrons in an atom to predict if the atom will form bonds.

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### CRITICAL THINKING

5. **Making Inferences** How can an atom that has 5 valence electrons achieve a full set of valence electrons?

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Section Review *continued*

6. **Applying Concepts** Identify the number of valence electrons in a barium atom.

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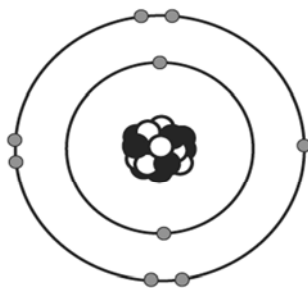
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**INTERPRETING GRAPHICS**

7. Look at the model below. How many valence electrons are in a fluorine atom? Will fluorine atoms form bonds? Explain.



Fluorine

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