

Section Review

Ionic Bonds

USING KEY TERMS

1. Use the following terms in the same sentence: *ion* and *ionic bond*.

2. In your own words, write a definition for the term *crystal lattice*.

UNDERSTANDING KEY IDEAS

- _____ 3. Which types of atoms usually become negative ions?

- a. metals
- b. nonmetals
- c. noble gases
- d. All of the above

4. How does an atom become a positive ion? a negative ion?

5. What are two properties of ionic compounds?

MATH SKILLS

6. What is the charge of an ion that has 12 protons and 10 electrons? Write the ion's symbol. Show your work below.

Section Review *continued*

CRITICAL THINKING

7. **Applying Concepts** Which group of elements gains two valence electrons when the atoms form ionic bonds?

8. **Identifying Relationships** Explain why ionic compounds are neutral even though they are made up of charged particles.

9. **Making Comparisons** Compare the formation of positive ions with the formation of negative ions in terms of energy changes.
